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so many fake sites. this is the first one which worked! Many thanks

Question 1.1:

Calculate the molecular mass of the following:

(i)  $H_2O$  (ii)  $CO_2$  (iii)  $CH_4$

Answer:

(i)  $H_2O$ :

The molecular mass of water,  $H_2O$

$= (2 \times \text{Atomic mass of hydrogen}) + (1 \times \text{Atomic mass of oxygen})$

$= [2(1.0084) + 1(16.00 \text{ u})]$

$= 2.016 \text{ u} + 16.00 \text{ u}$

$= 18.016$

$= 18.02 \text{ u}$

(ii)  $CO_2$ :

The molecular mass of carbon dioxide,  $CO_2$

$= (1 \times \text{Atomic mass of carbon}) + (2 \times \text{Atomic mass of oxygen})$

$= [1(12.011 \text{ u}) + 2(16.00 \text{ u})]$

$= 12.011 \text{ u} + 32.00 \text{ u}$

$= 44.01 \text{ u}$

(iii)  $CH_4$ :

The molecular mass of methane,  $CH_4$

$= (1 \times \text{Atomic mass of carbon}) + (4 \times \text{Atomic mass of hydrogen})$

$= [1(12.011 \text{ u}) + 4(1.008 \text{ u})]$

$= 12.011 \text{ u} + 4.032 \text{ u}$

$= 16.043 \text{ u}$

Question 1.2:

Calculate the mass percent of different elements present in sodium sulphate ( $Na_2SO_4$ ).

Answer:

The molecular formula of sodium sulphate is  $Na_2SO_4$ .

Molar mass of  $Na_2SO_4 = [(2 \times 23.0) + (32.066) + 4(16.00)]$

$= 142.066 \text{ g}$

Mass percent of an element  $= \frac{\text{Mass of that element in the compound}}{\text{Molar mass of the compound}} \times 100$

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