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Cool! I'am really happy

#Markus Jensen



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My friends are so mad that they do not know how I have all the high quality ebook which they do not!

#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

$N_2 + 3H_2 \rightarrow 2NH_3$
if the rate of reaction of consuming of N_2 is 0.25 mol per second ,
find the rate of reaction of forming NH_3 ?

formula for this reaction

$$-\frac{d}{dt} N_2 = -\frac{1}{3} \frac{d}{dt} H_2 = \frac{1}{2} \frac{d}{dt} NH_3$$

put the value there we get

$$0.25 = \frac{1}{2} \frac{d}{dt} NH_3$$

multiply by 2 there we get

$$0.5 = d/dt NH_3$$

answer is 0.5 M /sec

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